SC17a Group 1

Word	Pronunciation	Meaning
alkali metals		A group of very reactive metals found in group 1 of the periodic table.
group (chemistry)		A vertical column of elements in the periodic table. Elements in the same group generally have similar properties.
periodic table		The chart in which the elements are arranged in order of increasing atomic number.
reactivity		A description of how quickly or vigorously something reacts.

SC17b Group 7

Word	Pronunciation	Meaning
bleach		To take the colour out of something.
diatomic		Two atoms chemically bonded together.
disinfectant		Something that destroys or neutralises disease-carrying microorganisms.
halide		A compound formed between a halogen and another element such as a metal or hydrogen.
halogen		An element in group 7 of the periodic table.
salt		A compound formed by neutralisation of an acid by a base.

SC17c Halogen reactivity

Word	Pronunciation	Meaning
displacement reaction		When a more reactive element displaces a less reactive element from one of its compounds.
oxidation	*ox-id-**day**-shun*	A reaction in which oxygen is added to a substance loss of electrons by an atom or negative ion.
oxidised	***ox**-id-eyes'd*	When a substance has gained oxygen (or lost electrons) in a reaction.
redox		A reaction in which both oxidation and reduction occur.
reduced	*red-**yoos**'d*	When a substance has lost oxygen (or gained electrons) in a reaction.
reduction	*red-**duck**-shun*	A reaction in which oxygen is removed from a substance; gain of electrons by an atom or positive ion.

SC17d Group 0

Word	Pronunciation	Meaning
inert		Does not react.
noble gas	***nO**-bul **gas***	An unreactive gas in group 0 of the periodic table.

SC18a Rates of reaction

Word	Pronunciation	Meaning
product		A substance formed in a reaction.
rate		How quickly something happens.
reactant		A substance used up in a chemical reaction.
variable	***vair**-ee-ab-el*	A factor that can change.

SC18b Factors affecting reaction rates

Word	Pronunciation	Meaning
activation energy		The minimum amount of energy needed to start a reaction.
endothermic		A type of reaction in which energy from the surroundings is transferred to the products. The products have more stored energy than the reactants have.
exothermic		A type of reaction in which energy is transferred to the surroundings from the reactants. The products have less stored energy than the reactants have.

SC18c Catalysts and activation energy

Word	Pronunciation	Meaning
active site		The space in an enzyme where the substrate fits during an enzyme-catalysed reaction.
catalyst		A substance that increases the rate of a reaction without itself being used up.
denatured		An enzyme in which the shape of the active site has changed so much that its substrate no longer fits and the reaction can no longer happen.
enzyme		A protein produced by living organisms that acts as a catalyst to increase the rate of a reaction.
protein		A polymer made up of amino acids.
reaction profile		A diagram to show how the energy stored in substances changes during a chemical reaction.
substrate		A substance that is changed during a reaction.

SC19a Exothermic and endothermic reactions

Word	Pronunciation	Meaning
displacement reaction		A reaction where a more reactive element takes the place of a less reactive element in a compound.
endothermic	*end-O-**ther**-mik*	A reaction in which energy is transferred to the reactants from the surroundings. The temperature of the surroundings decreases.
exothermic	*ex-O-**ther**-mik*	A reaction in which energy is transferred from the reactants to the surroundings. The temperature of the surroundings increases.
neutralisation	*new-tral-i-**zay**-shun*	A reaction in which an acid reacts with an alkali or a base to produce a salt and water only.
precipitation		A reaction in which an insoluble product is formed from soluble reactants.
reaction profile		A diagram to show how the energy stored in substances changes during the course of a chemical reaction.

SC19b Energy changes in reactions

Word	Pronunciation	Meaning
activation energy		The minimum amount of energy needed by colliding particles for a reaction to happen.
bond energy		Energy needed to break one mole of a specified covalent bond.
covalent bond		A bond formed when a pair of electrons is shared between two atoms.
mole		A mole of something is 6.02×10^{23} particles of it. The mass of a mole of a substance is the relative formula mass expressed in grams.
reaction profile		A diagram to show how the energy stored in substances changes during the course of a chemical reaction.